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| **Alternative Evidence Gathering Template – Internal Assessment** |  |
| These templates must only be used to record student achievement and report results where remote assessment is the only practical option and the collection of direct assessment evidence from students has not been at all possible. ‘Alternative Evidence’ is student evidence for internally assessed standards that has been seen or heard within the teaching and learning programme. These templates do not signal a reduction in what is accepted for each grade, but rather a means of summarising evidence for reporting. These templates must be viewed in conjunction with the standard and assessment advice forwarded to schools to ensure that valid, credible and reliable assessment and learning has occurred before the standard is awarded. While physical evidence of student work does not need to be attached, the assessor decisions made must also be verified internally before reporting results. |  |
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| Student ID | Student 1 | Subject | Chemistry | Level | 2 |
| Notes |  | Standard No. | 91167 | Version | 2 |
| Standard Title | Demonstrate understanding of oxidation-reduction  | Credits | 3 |
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| **Achieved** | **Merit** | **Excellence** |
| Demonstrate understanding of oxidation-reduction.  | Demonstrate in-depth understanding of oxidation-reduction.  | Demonstrate comprehensive understanding of oxidation-reduction.  |
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| **Key requirements (list):** | A | M | E | **Describe or attach the evidence considered.**  | **Explain how the judgement was made.** |
| Identify, by name or formula, oxidation-reduction species.  |[ ] [ ] [ ]   |  |
| Describe each process as either oxidation or reduction with reasons – loss/gain of electrons (can be found in half equation) or the change in oxidation number.  |[ ] [ ] [ ]   |  |
| Writing correctly balanced half-equations (for oxidation and reduction).   |  |[ ] [ ]   |  |
| Linking these equations to the observations (reactants and products) and oxidation-reduction reasons (number of electrons lost/gained or the change in oxidation number assigned to the relevant element).  |  |[x] [ ]   |  |
| Justify oxidation and reduction by linking balanced half-equations, a correct overall equation, observations and electron transfer or change in oxidation number.  |  |  |[ ]   |  |
|  |  |  |  |  |  |
| **Sufficiency statement** | **Internal Verification**  |
| Achievement | All of A is required [x]  | Assessor: Date:  |
| Merit | All of A and M is required [x]  | Verifier: Date:  |
| Excellence | All of A, M and E is required [x]  | Verifier’s school:  |
| MARK OVERALL GRADE | N [ ]  | A [ ]  | M [ ]  | E [ ]  | Comments:  |

For the purpose of national external moderation:

* only six WORD templates are required where available
* samples are not required to be randomly selected
* there should be one each of N, A, M, E and up to 2 others
* descriptions of evidence and explanations of judgements are not required for all other students, and a spreadsheet may be used.