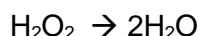


Electrochemical Cell**Reduction at the cathode:**

The reaction remains colourless.

- ① This is a reduction reaction as the oxidation number of O changes from -1 in H_2O_2 to -2 in H_2O

Oxidation at the anode:

This is an oxidation reaction as the oxidation number of Fe increases from +2 in Fe^{2+} to +3 in Fe^{3+} . An increase in oxidation number corresponds to oxidation.

Each Fe^{2+} loses 1 electron, losing electrons corresponds to oxidation.

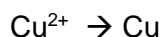
 E°

- ③ E° for H_2O_2 is the most positive so it will be reduced. This reaction creates electrical energy.

Electrolytic cell

From observations: The electrolysis of molten copper oxide produces bubbles of a colourless gas at one electrode and an orange solid is deposited at the other electrode.

Colourless gas is oxygen at the positive anode. Orange solid is copper and it is deposited at the negative cathode.

Reduction at the cathode (negative electrode)

The oxidation number of Cu goes from +2 in Cu^{2+} to 0 in Cu.

Each Cu^{2+} gains 2 electrons.

Oxidation at the anode (positive electrode)

The bubbles of colourless gas are oxygen.

Each O^{2-} loses 2 electrons.

- ② This process requires external electrical energy.